Additions and Corrections

NOTICE TO READERS.—For the convenience of those who may wish to cut out the corrections and attach them to the margins of the articles corrected, they have been printed only upon one side of the page.

VOL. 52, 1930

The Arsonation of Aromatic Aldehydes. By Albert B. Scott and Cliff S. Hamilton.

Pages 4122 ff. "In the article..., solutions of aromatic arsonic acids were warmed with hypophosphorous acid, the characteristic precipitation used as qualitative evidence of arsonation, and this statement occurs: 'When a solution of arsenic trioxide in dilute hydrochloric acid was treated with hypophosphorous acid under these conditions, no precipitate was obtained.' I find that a brown precipitate is formed when arsenic trioxide is warmed with hypophosphorous acid under the conditions described. The original statement is not correct."—ALBERT B. SCOTT.

VOL. 54, 1932

The Heat Capacity and Entropy of Carbon Monoxide. Heat of Vaporization. Vapor Pressure of Solid and Liquid. Free Energy to 5000°K. from Spectroscopic Data. By J. O. Clayton and W. F. Giauque.

Pages 2610 ff. The authors call attention to a small error in calculation and submit corrected data: "In our recent tabulation of the free energy function for carbon monoxide, a small systematic error occurred in the calculation. Also several of the signs were published incorrectly in the equation for rotational energy. Although the error (maximum about 0.04 cal./mole per deg.) is well within the limit of those contributed to the free energy of formation by the heat capacity and combustion data on graphite, it seems desirable to correct the table for future use in reactions where the other data are known with higher accuracy.

$$\epsilon_{\text{rotation}} = \left[1.853 - 0.020\left(v + \frac{1}{2}\right)\right]m^2 - \left[5.418 \times 10^{-6} - 6.918 \times 10^{-8}\left(v + \frac{1}{2}\right)\right]m^4$$

CORRECTED TABLE OF $(F^\circ - E_0^\circ)/T$ FOR CARBON MONOXIDE

T	$-(F^\circ-E_0^\circ)/T$	T	$-(F^\circ-E_0^\circ)/T$	T	$-(F^\circ-E_0^\circ)/T$	T	$-(F^\circ-E_0^\circ)/T$
250	**39.140**	950	48.503	2100	54.476	3600	58.830
298.1	**40.364**	1000	**48.876**	2200	54.843	3700	59.057
300	**40.408**	1050	49.230	2300	55.196	3800	59.278
350	41.479	1100	49.570	2400	55.534	3900	59.495
400	**42.408**	1150	49.896	2500	**55.860**	4000	**59.706**
450	43.228	1200	50.210	2600	56.175	4100	59.903
500	**43.963**	1250	50.513	2700	56.479	4200	60.114
550	44.629	1300	50.804	2800	56.773	4300	60.312
600	45.238	1400	51.359	2900	57.057	4400	60.505
650	45.801	1500	**51.880**	3000	**57.333**	4500	**60.695**
700	46.323	1600	52 370	3100	57 800	4600	60.879
750	**46.813**	1700	52.834	3200	57.860	4700	61.061
800	47.271	1800	53.275	3300	58.112	4800	61.239
850	47.703	1900	53.695	3400	58.358	4900	61.415
900	48.114	2000	**54.095**	3500	**58.597**	5000	**61.586**

The values given in heavy type have been directly calculated. The others interpolated.